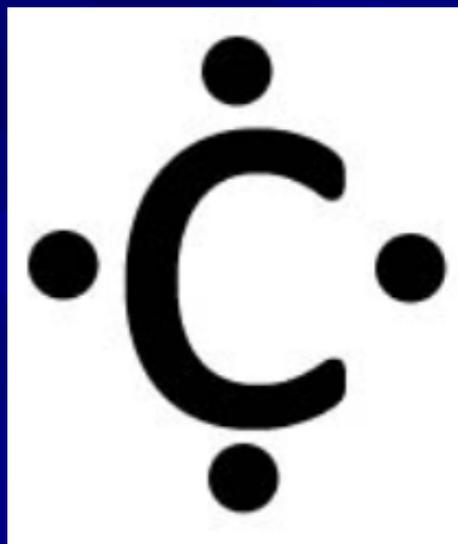


THE LEWIS NOTATION



- The Lewis notation is a representation of the atom in which the valence electrons are illustrated by DOTS placed around the *chemical symbol* of the element.

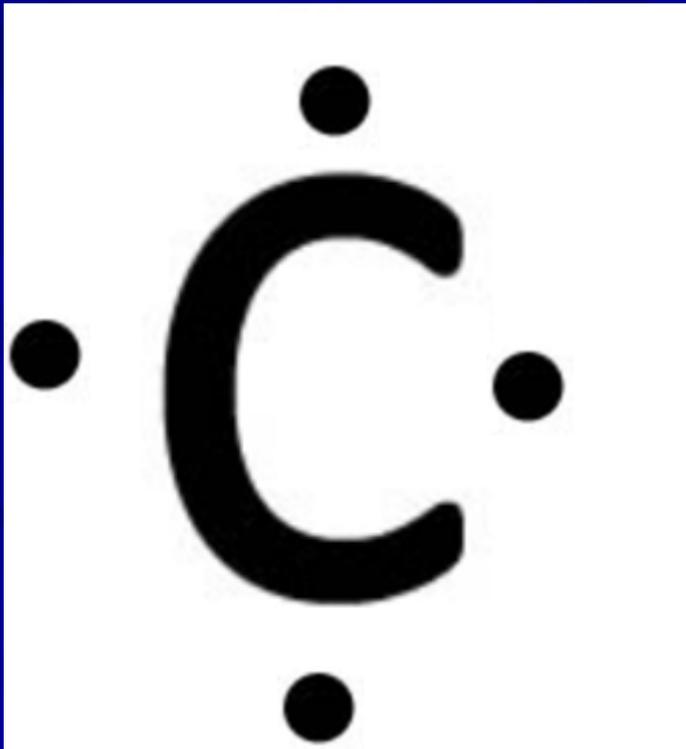
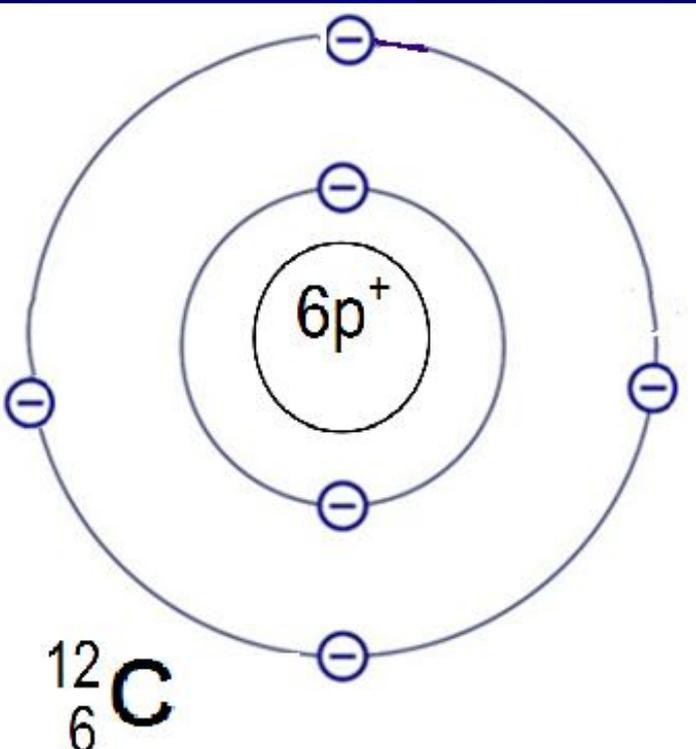


■ The electrons are placed one by one around the symbol, like the four *POINTS OF A COMPASS*

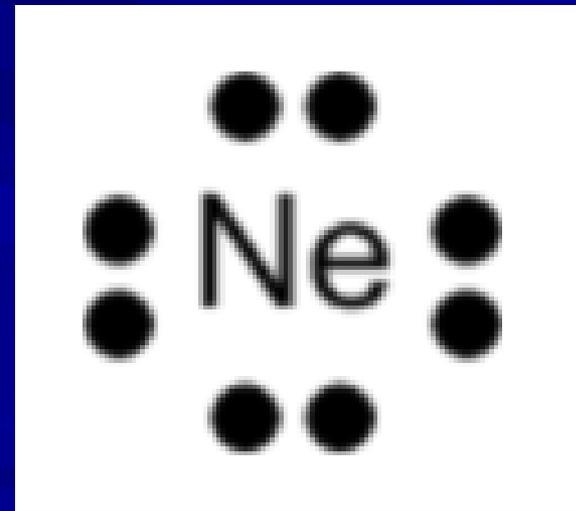
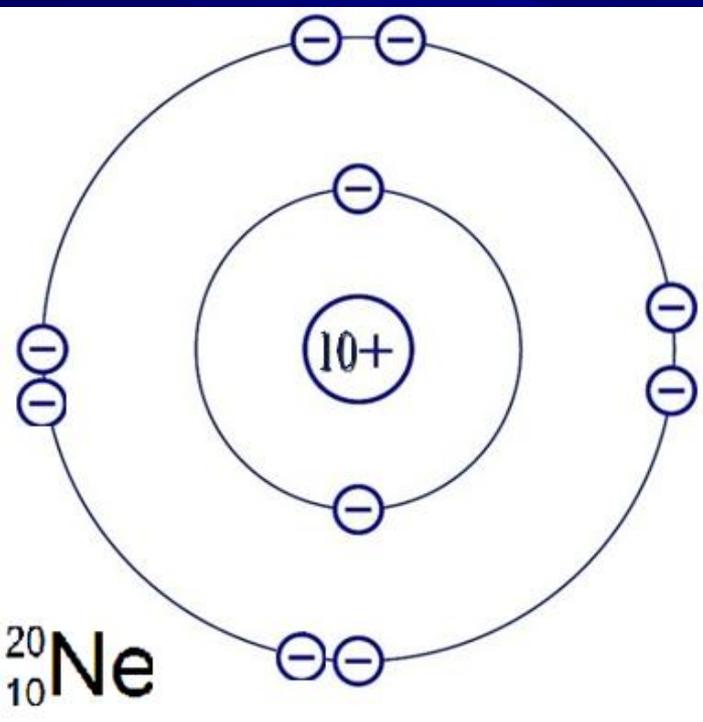
• Rutherford-Bohr DIAGRAM

■ EX: carbon:

LEWIS NOTATION



- If an element has more than four valence electrons, they are *DOUBLED* to form pairs of dots *AROUND* the element (for example, neon is represented by eight valence electrons).



Group

	1A 1	
1	1 H 1.00794 Hydrogen	2A 2
2	3 Li 6.941 Lithium	4 Be 9.01218 Beryllium



	3A 13	4A 14	5A 15	6A 16	7A 17	8A 18
	5 B 10.811 Boron	6 C 12.0107 Carbon	7 N 14.0067 Nitrogen	8 O 15.9994 Oxygen	9 F 18.9984 Fluorine	2 He 4.00260 Helium
						10 Ne 20.1797 Neon

➤ Fluorine:

➤ Lewis Notation:

